

Can $[M(H)_2(H_2)(PXP)]$ Pincer Complexes ($M = Fe, Ru, Os$; $X = N, O, S$) Serve as Catalyst Lead Structures for NH_3 Synthesis from N_2 and H_2 ?

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Abstract: The potential of pincer complexes $[M(H)_2(H_2)(PXP)]$ ($M = Fe, Ru, Os$; $X = N, O, S$) to coordinate, activate, and thus catalyze the reaction of N_2 with classical or nonclassical hydrogen centers present at the metal center, with the aim of forming NH_3 with H_2 as the only other reagent, was explored by means of DF (density functional) calculations. Screening of various complexes for their ability to perform initial hydrogen transfer to coordinated N_2 showed ruthenium pincer complexes to be more promising than the corresponding iron and osmium analogues. The ligand backbone influences the reaction dramatically: the presence of pyridine and thioether groups as backbones in the ligand result in inactive

catalysts, whereas ether groups such as γ -pyran and furan enable the reaction and result in unprecedented low activation barriers (23.7 and 22.1 kcal mol⁻¹, respectively), low enough to be interesting for practical application. Catalytic cycles were calculated for $[Ru(H)_2(H_2)(POP)]$ catalysts (POP = 2,5-bis(dimethylphosphanyl)methylfuran and 2,6-bis(dimethylphosphanyl)methyl- γ -pyran). The height of activation barriers for the furan system is somewhat more advantageous. Forma-

tion of inactive metal nitrides has not been observed. SCRF calculations were used to introduce solvent (toluene) effects. The Gibbs free energies of activation of the numerous single reaction steps do not change significantly when solvent is included. The reaction steps associated with the formation of the active catalyst from precursors $[M(H)_2(H_2)(PXP)]$ were also calculated. The otherwise inactive pyridine ligand system allows for the generation of the active catalyst species, whereas the ether ligand systems show activation barriers that could prohibit practical application. Consequently the generation of the active catalyst species needs to be addressed in further studies.

Keywords: activation energy • ammonia • density functional calculations • homogeneous catalysis • pincer complexes • ruthenium

Introduction

Although heterogeneously catalyzed ammonia synthesis has long been known and is well understood,^[1] the homogeneously catalyzed synthesis of NH_3 from N_2 and H_2 by a defined transition-metal complex is one of the unresolved challenges in chemistry.^[2] Significant efforts, many of them driven by an attempt to understand biological nitrogen reduction, have been made to synthesize and characterize transition-metal complexes that bond with and activate the

N_2 molecule. A wealth of metal complexes (for example, titanium,^[3] zirconium,^[4,9c] vanadium,^[5] molybdenum,^[6] and iron complexes^[7]) were shown to bond with and activate the N_2 molecule to varying extents, some of them even breaking the N–N structure motif completely. Many of the theoretical and spectroscopic data available have been analyzed.^[8] Most of the synthetically derived transition-metal complexes tested for catalytic reduction of N_2 to form ammonia proved inactive. However, very stimulating progress has been made recently with regard to both reduction methods (N_2/H_2 and N_2/H^+ /reducing agent). For instance, in the former case the complex $[[Zr(\eta^5-C_5Me_4H)_2(\mu^2, \eta^2, \eta^2-N_2)]]$ with side-on-bonded N_2 reacted in the presence of H_2 to yield NH_3 . However, the starting complex was consumed, rendering the reaction stoichiometric.^[9a] These experimental results were rationalized on the basis of DFT calculations.^[9b–e]

With regard to mimicking the biological process of N_2 fixation and reduction to NH_3 using protons and reducing agents, in a seminal work by Schrock and co-workers the synthetic complex $[Mo(N_2)(HIPTN_3N)]$ ($HIPTN_3N =$ hexa-

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isopropylterphenyltriamidoamine) proved to be an active catalyst for catalytic NH_3 synthesis.^[10] In this approach N_2 was first coordinated at the molybdenum complex, and subsequently protonated/reduced repeatedly in the presence of acids (proton source) and metallocenes (reducing agent), generating NH_3 . Meanwhile some intermediates were characterized spectroscopically and crystallographically, while important details of the mechanism were elucidated by computational studies.^[11]

In a very recent DFT study on molybdenum, ruthenium, and osmium model complexes closely related to the original Schrock catalyst, we investigated the catalytic potential of metal hydrides of general formula $[\text{M}=\text{H}(\text{tren})]$ (tren = tris(2-amidoethyl)amine or tris(2-amidoethyl)phosphane) for NH_3 formation using only N_2 and H_2 .^[12] For these complexes with formal oxidation state +4 at the metal center, catalytic cycles could be calculated. For the molybdenum system we found activation barriers that were much too high to consider any experimental approach. We also observed that for the ruthenium and osmium systems some of the many single reaction steps have much lower activation barriers than the molybdenum system; these seem to be related to the higher electron count of Ru and Os. Despite

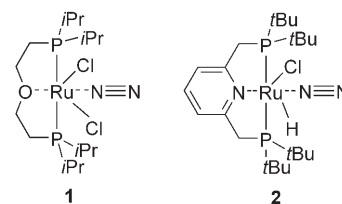


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this encouraging result for a variety of single transformations, activation energies well above 30 kcal mol^{-1} were calculated, indicating a need for further theoretical investigations. We reasoned that a computational study focusing on Group 8 metal complexes in the oxidation state +2 should help to identify catalyst lead structures for the reduction of N_2 by H_2 . In this study we investigated iron, ruthenium, and osmium complexes of the general type **A** (D = donor

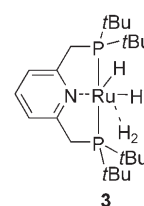
group) with various pincer ligands, for the following reasons:

- 1) We assume that octahedral coordination polyhedra will be the dominating coordination geometry for this type of complex (with Ru^{2+} as the central atom, for example). Accordingly, the pincer ligand should occupy three of the six coordination sites permanently and thus help to shield the metal center from the formation of—presumably—unreactive dimeric species containing, for example, the M-N-N-M structural motif. The formation of the M-N-N-M unit was proven by crystal structure analyses for $[\{\text{RuH}(\text{pcp})\}_2(\mu\text{-N}_2)]$,^[13] $[\{\text{RuH}(\text{N}_2)(\text{pcp})\}_2(\mu\text{-N}_2)]$ (pcp = 1,3-(CH_2PrBu_2) $_2\text{C}_6\text{H}_4$),^[13] and $[\{\text{RuH}_2(\text{N}_2)(\text{PiPr}_3)_2\}_2(\mu\text{-N}_2)]$ ^[14] as well as for $[\{\text{RuCp}(\text{dippe})\}_2(\mu\text{-N}_2)]^{2+}$ (Cp = cyclopentadiene, dippe = 1,2-bis(diisopropylphosphanyl)ethane) and $[\{\text{RuCp}(\text{PET}_3)_2\}_2(\mu\text{-N}_2)]^{2+}$.^[15] In practical applications, appropriate shielding preventing dimer formation can be introduced into the model complexes studied here by attaching sterically demanding substituents (R) to the P centers of the ligand. The decision on which substituent is appropriate is not addressed in this study.
- 2) Mononuclear ruthenium complexes, such as **1**^[16] and **2**,^[17] with an ether and a nitrogen donor group, respectively,



in the ligand backbone and each bonding one N_2 molecule end-on, had already been shown to be available synthetically and were characterized crystallographically, thus proving that N_2 complexes of this kind are accessible. For **1** it was also shown that N_2 can be displaced by H_2 , indicating that dissociation or reassociation of either H_2 or N_2 should be possible under appropriate conditions.

- 3) Nonclassical ruthenium complexes such as **3**, which among other complexes of this type are currently under investigation by our group, are able to coordinate or dissociate H_2 molecules reversibly and also to activate aromatic C-H bonds under mild conditions, proving a significant catalytic activity.^[18]



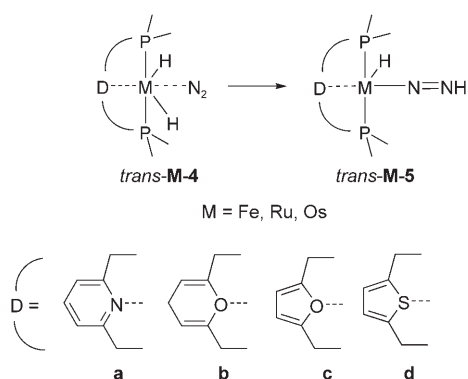
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- 4) Iron and osmium complexes are included in parts of the study to obtain information about their potential catalytic performance (that is, existence of local minima and transition states and relative energies) in one of the crucial steps of nitrogen reduction.

Results and Discussion

This work is in three parts: 1) calculation of initial transfer of one hydrogen center onto coordinated N_2 ; 2) calculation of complete catalytic cycles for N_2 reduction (sometimes with the inclusion of solvent); 3) calculation of local minima and transition states necessary to convert the catalyst precursors into active catalyst species.

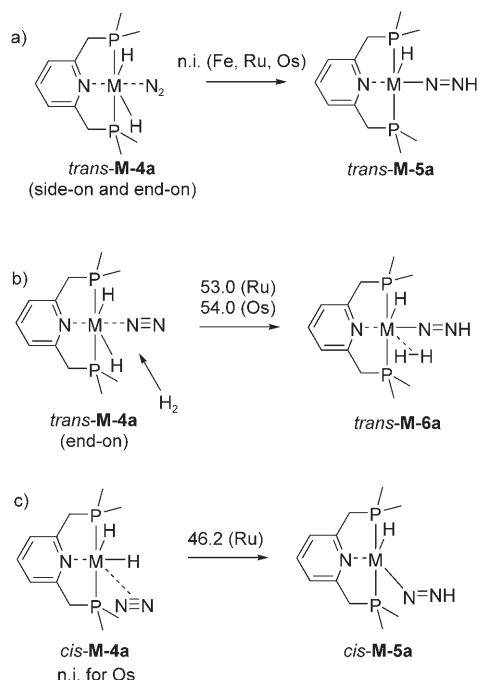
Initial hydrogen transfer: As the aim of the study was calculation of complete catalytic cycles for N_2 reduction, it was efficient to concentrate on the initial hydrogen transfer as a first part of this computational catalyst screening. As our previous theoretical work^[12] in accordance with conventional wisdom had shown that one of the energetically very demanding reaction steps is the initial transfer of a metal-bonded hydride center onto coordinated N_2 , we initiated our study by using hypothetical structures **4** and **5**, which are model complexes derived from catalyst **3**. We calculated local minima and transition states associated with the initial transfer of one metal-bonded hydride onto coordinated N_2 . For the bisphosphane ligands with ether groups in the backbone, we then calculated catalytic cycles for N_2 reduction. Scheme 1 shows the reaction investigated together with the



Scheme 1. Computationally investigated reaction of *trans-M-4* to *trans-M-5* with a variety of metals and ligand backbones.

ligand backbones used. In this way we hoped to assess the influence of different metals and different donor abilities of groups D in the ligand backbone.

Ligand backbone **a** was investigated first. Complex *trans-Ru-4a* could be identified with N₂ both in side-on and in end-on coordination mode, but to our disappointment it was not possible to locate a transition state leading to *trans-Ru-5a* (Scheme 2a). Similar results were obtained for the corresponding Fe and Os complexes, indicating that *trans-M-4a* is not a suitable catalyst candidate for N₂ reduction when a metal-bonded hydride is to be transferred (Scheme 2a). Alternatively *trans-Ru-4a* can react with one additional H₂ molecule, which approaches the complex from the outer co-

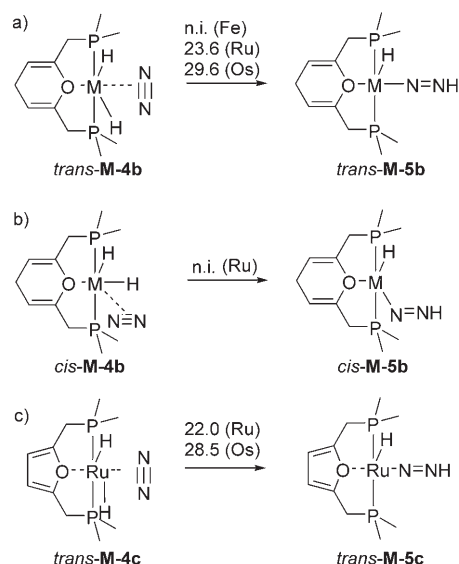


Scheme 2. Calculated PNP-M complexes (M=Fe, Ru, Os) involved in initial hydrogen transfer to N₂ with activation energies ΔE^\ddagger (kcal mol⁻¹) shown above the reaction arrows. For details, see the text; n.i.=not identified. Structural representations with selected parameters are given in the Supporting Information.

ordination sphere by breaking the H-H bond and transferring one H atom to the N₂ moiety, while the other forms a nonclassically bonded H₂ molecule with the hydride center present at the metal center. In this way complex *trans-Ru-6a* is generated (Scheme 2b). The activation energy for this step is 53.0 kcal mol⁻¹. All subsequent reaction steps aiming at the formation of NH₃ starting with catalyst *trans-Ru-4a* would rely on this initial reaction and consequently prohibit practical application. The calculations for the corresponding osmium complexes indicated the activation energy to be similar (54.0 kcal mol⁻¹).

However, using *cis-Ru-4a* with N₂ bonded side-on as the starting complex, generation of a metal-bonded M=N=NH unit is possible. However, the activation energy for this reaction of 46.2 kcal mol⁻¹ indicates the reaction is not possible in practice (Scheme 2c). All computationally derived results for ligand system **a** are in accord with our repeated experimental attempts to produce NH₃ in the presence of N₂, H₂, and catalyst precursor [Ru(H)₂(H₂)(PXP)] (PXP=bis(2,5-di-*tert*-butylphosphanyl)methylpyridine), in which no formation of NH₃ occurred.^[19] The corresponding osmium complex *cis-Os-4a* with N₂ bonded side-on in the starting material could not be identified.

In the next step of the study γ -pyran (ligand backbone **b**) was exchanged for the pyridine ring, as we wanted to introduce a less electron-donating ligand with as little structural change as possible. Such POP-Ru pincer complexes with this ligand are not yet known, to the best of our knowledge. In contrast to the corresponding pyridine complex, the γ -pyran compound enables the H transfer from the ruthenium center to the N atom of the N₂ unit, as we identified reactant *trans-Ru-4b*, product *trans-Ru-5b*, and transition state **TS(trans-Ru-4b/5b)** (Scheme 3a).



Scheme 3. Calculated POP-M complexes (M=Fe, Ru, Os) involved in initial hydrogen transfer to N₂ with activation energies ΔE^\ddagger (kcal mol⁻¹) shown above the reaction arrows. For details, see the text; n.i.=not identified. Structural representations with selected parameters are given in the Supporting Information.

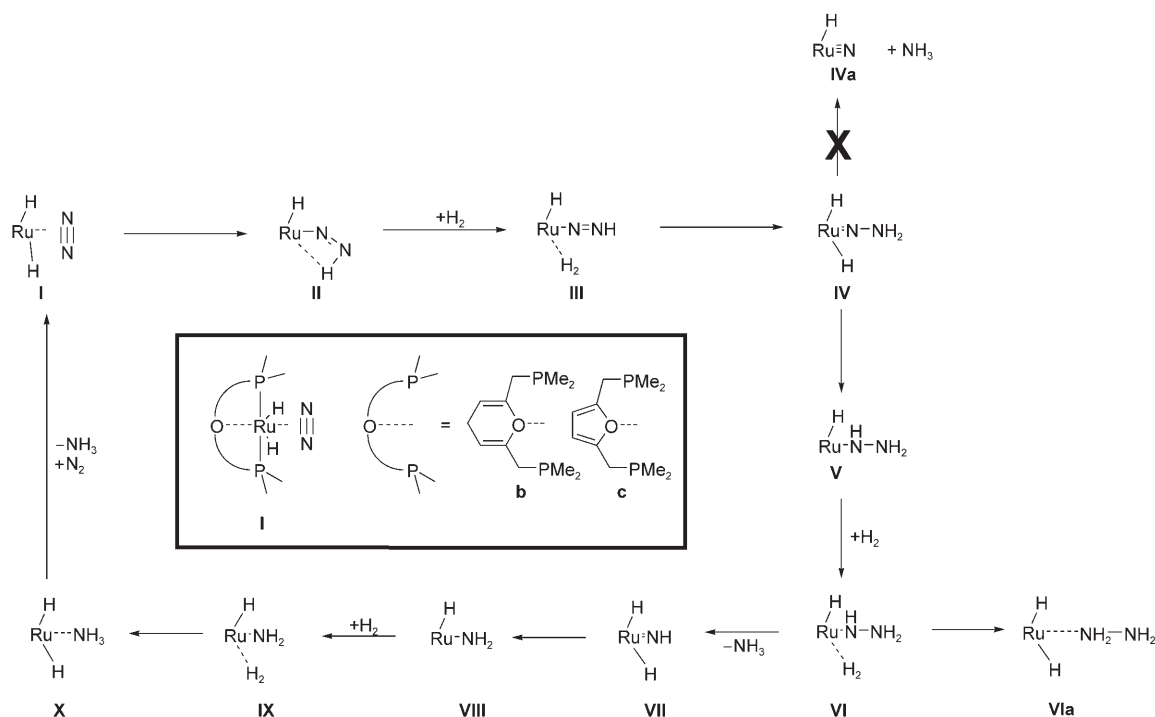
To our surprise the activation energy ΔE^\ddagger was fairly low ($23.6 \text{ kcal mol}^{-1}$). The corresponding iron complex *trans-Fe-4b* was identified as well as the product *trans-Fe-5b*, but the transition state **TS**(*trans-Fe-4b/5b*) does not exist. The corresponding osmium compounds were identified, but the activation barrier for the reaction from *trans-Os-4b* to *trans-Os-5b* was significantly higher ($\Delta E^\ddagger = 29.6 \text{ kcal mol}^{-1}$; Scheme 3a). For the ruthenium complex, calculations to check whether the same reaction would occur if the N_2 molecule was coordinated side-on and cis to the γ -pyran ring showed the starting complex *cis-Ru-4b* to exist, but no transition state could be identified. This indicates that the reaction is possible only when the N_2 molecule is coordinated *trans* to the γ -pyran ring. After furan was exchanged for the γ -pyran ring, *trans-Ru-4c*, *trans-Ru-5c* and **TS**(*trans-Ru-4c/5c*) were also identified, the activation energy being slightly lower than that of the γ -pyran system ($\Delta E^\ddagger = 22.0 \text{ kcal mol}^{-1}$; Scheme 3c). Again the corresponding osmium complexes were identified, but yielded a clearly higher activation barrier ($\Delta E^\ddagger = 28.5 \text{ kcal mol}^{-1}$). Interestingly, with the ether-containing ligand systems the activation energies were similar and slightly above 20 kcal mol^{-1} , whereas although the corresponding osmium complexes were identified, their activation barriers were approximately 6 kcal mol^{-1} higher. We therefore excluded any osmium compounds from the rest of the study. As the transition state for the γ -pyran system could not be identified using the corresponding iron complex, we ruled out iron at this point and did no calculations for any further iron complexes.

To check the influence of thioether groups, we did calculations for the corresponding thiophene complexes *trans-*

Ru-4d and *trans-Ru-5d* (ligand system **d**) and found both reactant and product to exist on the hypersurface. However, no transition state was found, indicating this donor not to be suitable for the generation of an active catalyst.

Calculation of catalytic cycles for N_2 reduction: The results obtained so far indicate a preference for ruthenium as the metal center rather than iron and osmium, while ligand backbones with ether groups seem to be favored over the stronger nitrogen- or sulfur-donor-containing ligands. We therefore calculated the catalytic cycles for the γ -pyran and the furan system to find out whether there were any closed catalytic cycles (Scheme 4). We aimed to compare the reaction energies of the two systems to estimate how much these two seemingly alike ligands would influence the different reaction steps; the corresponding energy profiles for the γ -pyran and the furan system are shown in Figure 1 (they are very similar with one exception, *vide infra*; the complexes have been renumbered, for clarity). Table 1 contains relative energies of the complexes associated with the reduction cycles.

Summarizing the reaction mechanisms briefly for both the γ -pyran and the furan system: the N_2 reduction starts with the first transfer of a hydride to the coordinated N_2 molecule (**I**→**II**). Subsequently a new H_2 molecule is coordinated at the metal center (**III**), which then reacts with the $\text{Ru}=\text{N}=\text{NH}$ moiety. The $\text{H}-\text{H}$ bond is broken and one hydrogen center is transferred to the external nitrogen atom resulting in the formation of a $\text{Ru}=\text{N}-\text{NH}_2$ unit, while the other hydrogen atom is bonded to the metal (**IV**). In principle one of the hydride centers could at this stage react with the NH_2



Scheme 4. Calculated catalytic cycle for POP-Ru pincer complexes with ligand backbones **b** and **c**.

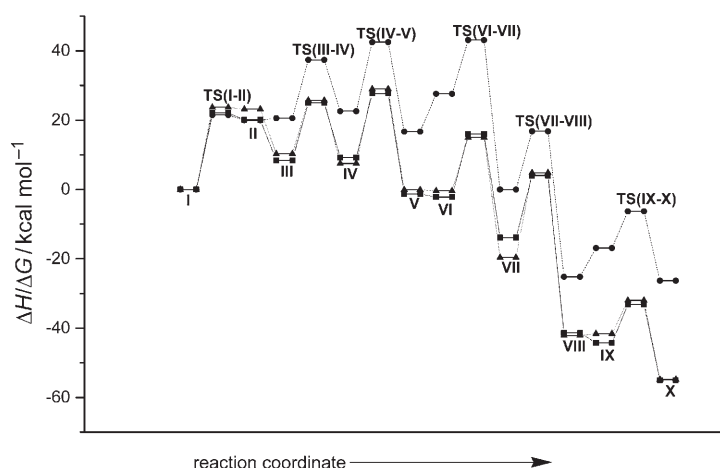


Figure 1. Energy profiles of the main part of the calculated catalytic cycles: gas-phase energies ΔE (kcal mol^{-1}) of the catalysts with \blacktriangle the γ -pyran (**b**) and \blacksquare the furan (**c**) ligand systems; \bullet Gibbs free energies ΔG (kcal mol^{-1}) of the catalyst with the furan ligand system (**c**) in toluene as the solvent.

Table 1. Relative energies ΔE and Gibbs free energies ΔG (all in kcal mol^{-1}) of complexes associated with N_2 reduction using ligand systems **b** and **c** in the gas phase (**b**, **c**) and in toluene (**c**).

	ΔE (b) gas	ΔE (c) gas	ΔG (c) toluene
I	0.00	0.00	0.00
II	23.2	20.1	20.0
III	10.3	8.4	20.6
IV	7.5	9.2	22.6
V	-0.1	-1.3	16.7
VI	-0.4	-2.2	27.6
VII	-19.6	-13.9	0.0
VIII	-42.1	-43.2	-25.2
IX	-42.6	-44.3	-16.9
X	-54.8	-55.0	-26.3
TS(I-II)	23.7	22.1	21.6
TS(III-IV)	25.7	25.0	37.4
TS(IV-V)	29.0	27.7	42.5
TS(VI-VII)	15.0	16.0	43.1
TS(VII-VIII)	4.8	4.0	16.8
TS(IX-X)	-32.0	-33.2	-6.3
TS(VI-VIa)	10.1	9.2	36.8

group, forming the first NH_3 molecule and the corresponding ruthenium nitride (**IVa**). Interestingly, although the product could be located on the hypersurface, the transition state does not exist, as repeated attempts to locate it failed. This is a clear contrast to the catalysts containing the tris-amidoamine ligand, which we reported earlier. For these complexes we found the corresponding transition state leading the reaction to the metal nitride. This has direct consequences for potential practical applications: here the impossibility of locating a transition state to generate **IVa** indicates that these types of pincer complexes would not form unreactive nitrides; in other words, in a potential reaction the catalyst would not be consumed.

In an alternative step the hydride center is transferred to the internal nitrogen atom, resulting in the formation of a Ru-NH-NH_2 unit and creating a vacant coordination site at

the metal (**V**), which is subsequently occupied by H_2 (**VI**). Now the reaction can result in either the formation of the first NH_3 molecule of the cycle by breaking the H-H bond of the coordinated H_2 molecule and transfer of one hydrogen to the NH_2 group (**VII**), or reaction of the hydrogen center with the nitrogen atom of the NH unit resulting in the formation of a coordinated hydrazine molecule (**VIa**). Kinetically the latter reaction ($\Delta E^\ddagger = 10.3$ and $9.9 \text{ kcal mol}^{-1}$ for ligand systems **b** and **c**, respectively) is clearly more favored than the former one ($\Delta E^\ddagger = 15.4$ and $16.3 \text{ kcal mol}^{-1}$ for ligand systems **b** and **c**, respectively), but hydrazine formation is endothermic. As the barrier for the reaction of **VIa** to **VI** is $23.9 \text{ kcal mol}^{-1}$ for the γ -pyran complex and $23.0 \text{ kcal mol}^{-1}$ for the furan complex, the hydrazine complexes **VIa** should both be able to re-enter the cycle. Alternatively, after the formation of NH_3 from the reaction of **VI** to **VII** an imido unit is present at the metal center that, by reaction with one of the hydride centers, is converted to an amido unit (**VIII**). This complex again has a vacant coordination site which is occupied by H_2 (**IX**) and in a final reaction the H-H bond is broken, while the second NH_3 molecule is created and coordinates to the metal center (**X**). By dissociation of NH_3 and re-association of N_2 the cycle can start again (for solvation effects, see below).

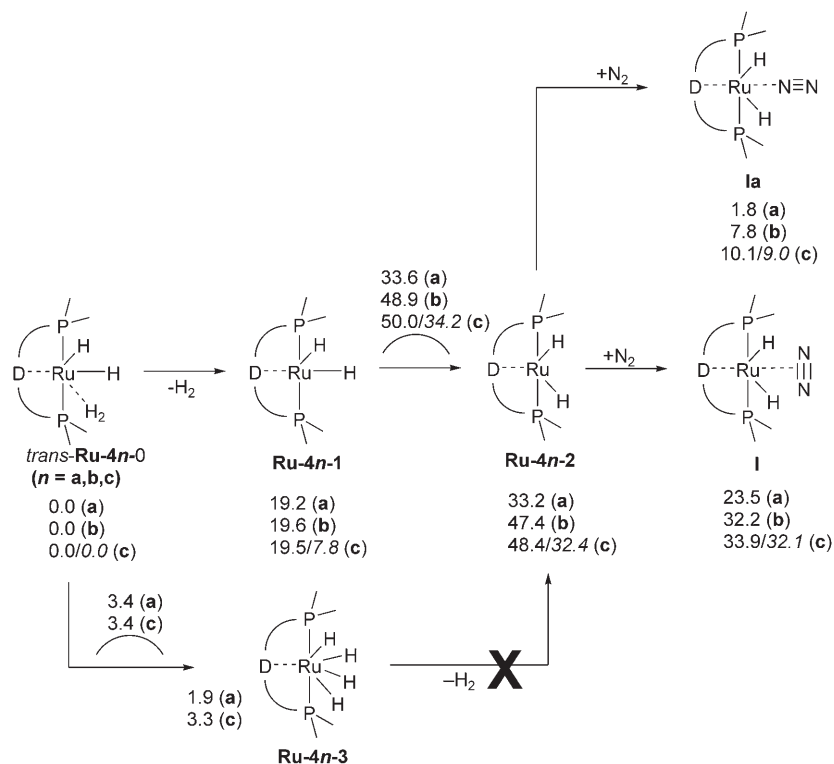
Most of the calculated activation barriers for the γ -pyran-containing catalyst system are below $20.0 \text{ kcal mol}^{-1}$, suggesting these steps are possible in practice. For the furan system all steps except for the initial H transfer to N_2 (vide supra) are below $20.0 \text{ kcal mol}^{-1}$, indicating this cycle to be preferred over the γ -pyran catalyst system. The most important difference in activation barriers between the two ligands occurs at the step in which the imido group ($\text{Ru}=\text{NH}$) reacts to form the amido group ($\text{Ru}-\text{NH}_2$; reaction **VII** \rightarrow **VIII**). In the γ -pyran system this step requires an activation energy ΔE^\ddagger of $24.4 \text{ kcal mol}^{-1}$, while in the furan system it needs only $16.6 \text{ kcal mol}^{-1}$. All other steps show much less pronounced differences.^[20a,b]

We re-optimized the geometries of the gas-phase structures for the furan-containing catalyst system in the presence of a relatively unpolar solvent (toluene), to investigate whether the inclusion of this reaction medium would alter structures and energies to a significant extent. We found no significant geometric changes in structures. Interestingly, the activation barriers ΔG^\ddagger in solution also do not change significantly relative to the gas-phase barriers. The highest activation barrier in solution was found, as in the gas phase, for the initial hydrogen transfer (reaction **I** \rightarrow **II**; $\Delta G^\ddagger = 21.5 \text{ kcal mol}^{-1}$). In the gas-phase addition of H_2 to **II** yielding **III** is practically thermoneutral, with a very small exothermic trend ($\Delta G = -0.4 \text{ kcal mol}^{-1}$). In solution the same applies, with the trend reversed ($\Delta G = 0.6 \text{ kcal mol}^{-1}$). Upon addition of H_2 in the gas phase, complex **V** reacts to **VI** in an endothermic reaction ($\Delta G = 9.8 \text{ kcal mol}^{-1}$), while in solution the Gibbs free energy is raised by $10.9 \text{ kcal mol}^{-1}$. An increase in energy of the same magnitude can be observed upon addition of H_2 to **VIII** yielding **IX**, as ΔG for the gas phase increases by $9.3 \text{ kcal mol}^{-1}$ and by $8.3 \text{ kcal mol}^{-1}$ for

the reaction in solution. Though endothermic, the addition of H₂ to these complexes should be possible as the energy increase is not too pronounced. To close the catalytic cycle the NH₃ molecule in **X** must dissociate from the complex while N₂ must be added to it. Transition states could not be located, neither for the dissociation of NH₃ from **X** nor for the addition of N₂, although several attempts were made. Single-point calculations for the dissociation of the NH₃ molecule from **X** to estimate the energy curve showed a nearly linear increase in energy upon removing NH₃ from the complex in 0.1 Å steps up to a distance of 4.150 Å between the metal and the nitrogen center (Ru–N in complex **X** = 2.152 Å). The same method for the addition of N₂ yielded a nearly linear decrease in energy when N₂ approaches the metal complex fragment from a distance of 4 Å. According to this result either it would not be possible to locate a transition state, or its barrier would be too small to be detected with the methods used here. As the free energy of solvation of ammonia ($\Delta G_{\text{solv}} = -6.56 \text{ kcal mol}^{-1}$) indicates that NH₃ dissolves much better in toluene than N₂ ($\Delta G_{\text{solv}} = 0.18 \text{ kcal mol}^{-1}$), it is reasonable to assume that the NH₃ molecule can dissociate from complex **X** and instead the complex fragment coordinates with N₂, in this way regenerating **I**. In fact for the reaction of **X** with N₂ to **I** and NH₃ ΔG in solution amounts to 13.6 kcal mol⁻¹. This is an endothermic step, but the energy increase is small enough to be reached by experimental systems.

Generation of the active catalyst species: The catalytic cycles obtained for the γ -pyran and furan ligand systems can only be active when it is possible to generate the starting complex **I** from the corresponding hydrides *trans*-**Ru-4b** and *trans*-**Ru-4c**. We therefore calculated the local minima and transition states associated with the generation of the active catalyst **I** from the starting materials (for comparison we included the otherwise inactive *trans*-**Ru-4a**, that is, the pyridine ligand system). In Scheme 5 the corresponding structures and relative energies are shown.

There are two ways in which **I** can be generated from the parent metal hydrides. The nonclassically bonded H₂ molecule in *trans*-**Ru-4n-0** can dissociate from the complex, generating complex **Ru-4n-1** ($n = \text{a, b, c}$). Instead of occupying



Scheme 5. Structures and energies [ΔE , ΔE^\ddagger for ligand systems **a–c** and ΔG , ΔG^\ddagger for ligand system **c** in toluene solution (italics) (all in kcal mol⁻¹)] of complexes associated with the generation of the active catalyst species **I** from the catalyst precursor *trans*-**Ru-4n** ($n = \text{a, b, c}$). The activation barriers for the reaction of **Ru-4n-1** to **Ru-4n-2** and for *trans*-**Ru-4n** to **Ru-4n-3** are given above the curve over the reaction arrow.

the vacant coordination site with either N₂ or solvent, the hydride center *trans* to the ligand can rearrange to yield the corresponding *cis* complex **Ru-4n-2**. The rearrangement requires substantially different amounts of energy for the different ligand systems. With the pyridine-containing ligand system **a**, the reaction can be accomplished, as the activation energy amounts to 14.4 kcal mol⁻¹. With the γ -pyran system **b**, the activation energy is much higher (29.3 kcal mol⁻¹), as it is in the furan system **c** (27.8 kcal mol⁻¹), indicating that chances of accomplishing this reaction in practice might be low. Repetition of the calculations for ligand system **c** in the presence of solvent (toluene) indicated that the barrier is slightly lower ($\Delta G^\ddagger = 26.4 \text{ kcal mol}^{-1}$). An alternative could be the rearrangement of the H₂ unit of *trans*-**Ru-4n-0** forming **Ru-4n-3**, which is formally a Ru⁴⁺ complex as the central two hydrogen atoms are singly bonded to the metal, but not to each other. The activation energy required to form **Ru-4n-3** calculated for ligand systems **a** and **c**, 3.4 kcal mol⁻¹ in both cases, is easily possible to achieve. From this complex an H₂ molecule might theoretically form *trans* to the ligand backbone and dissociate from the complex, but as no transition state for such a reaction was identified it is considered improbable.

It is important with regard to potential practical uses of this type of catalyst that the end-on-bonded isomers **Ia** are significantly more stable for all the ligand systems **a, b**, and **c** than the corresponding side-on-bonded isomers **I**. This sta-

bility pattern is in accord with the experimentally obtained crystal structures of complexes **1** and **2**, in which N₂ is bonded end-on. However, according to IRC (intrinsic reaction coordinate) calculations, the calculated reaction can only proceed starting from the complexes with the N₂ molecule bonded side-on (**I**). Though it is the thermodynamically less stable isomer, it is not impossible that **I** forms in solution. As a consequence the reaction should generally be possible. From the activation energies associated with the generation of the active catalyst, it is evident that the catalyst backbone of the ether systems should be modified to lower the transition-state energy associated with the rearrangement from **Ru-4n-1** to **Ru-4n-2**. If this issue can be resolved, ruthenium pincer complexes might develop into a catalyst class with promising potential for ammonia synthesis.

Conclusion

In this work we have shown by means of DFT calculations that POP–Ru pincer complexes can be used to generate computationally derived closed catalytic cycles for the synthesis of NH₃ from N₂ and H₂ with appreciably low activation barriers. To the best of our knowledge these are the lowest activation barriers reported to date for molecular Ru²⁺-complex-catalyzed reduction of N₂ to NH₃ with H₂. In contrast, PNP pincer ligands (containing pyridine as N donor) as well as PSP ligands (containing thiophene as S donor) do not allow NH₃ synthesis, as the initial transfer of metal-bonded hydrogen to coordinated N₂ is not possible in the way found for the POP ligands. Alternative pathways involving H₂ bond metathesis of an additional H₂ molecule attacking the complex without bonding to the metal before reaction is possible for PNP–Ru complexes in general. However, practical application might be hindered due to the high activation barriers. Comparison of POP–Ru catalyst systems with POP–Os and POP–Fe catalysts hints at less favored (that is, higher) activation barriers or/and at nonexistent local minima and transition states, respectively, in this way clearly favoring POP–Ru catalysts. The inclusion of solvent effects for the furan-containing POP–Ru catalyst indicates Gibbs free activation energies to be very similar to the gas-phase activation barriers. To the best of our knowledge this is the first time a POP–Ru pincer complex has been shown to be potentially useful for homogeneously catalyzed NH₃ production from N₂ and H₂. However, issues with the generation of the active catalyst species from the catalyst precursor due to high activation barriers for rearrangement of metal-bonded hydride centers might impede practical realization of the reaction. Further theoretical work to clarify and improve this situation is under way.

Computational Details

All calculations were carried out using the Gaussian 03^[21] program series, implemented with the B3LYP^[22] hybrid functional. For nonmetal atoms

the TZVP basis set by Ahlrichs et al., as implemented in Gaussian 03,^[23] was used, while for the metal atoms a (311111/221111/411/1 (Fe); 311111/221111/411 (Ru, Os)) basis set was used in combination with the Stuttgart/Köln ECP.^[24] All stationary points were checked by frequency calculations to be either minima (*i*=0) or transition states (*i*=1), while IRC calculations carried out for most of the identified transition state structures proved that the correct products and reactants were connected by that transition state. Tables with energies, zero-point vibrational corrected energies, enthalpies, and Gibbs free energies (*T*=298 K, *p*=1 bar) are contained in the Supporting Information, as well as Cartesian coordinates and structural representations of most of the calculated compounds. Calculations (geometry optimization followed by frequency calculation) in the presence of a solvent (toluene, dichloromethane) were carried out using the self-consistent reaction field (SCRf) formalism, as implemented in Gaussian 03 employing the IEF-PCM (integral equation formulation of the polarizable continuum model)^[25] together with the united atom topological model for radii. Extra spheres for hydrogen were added for H₂, NH₃, N₂H₄, and all the hydrogen atoms in the complexes present at the Ru center and at the mono- or dinitrogen moieties in all the reaction steps. During the geometry optimizations and the subsequent frequency calculations, the calculation of dispersion solute–solvent interaction energy, of repulsion solute–solvent interaction energy, and of the cavitation energy was switched off.

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- [19] With DF calculations for [Ru(H)₂(H₂)(PXP)] (PXP = bis(2,5-di-*tert*-butyl-phosphanylmethyl)pyridine) as the catalyst we found the dissociation of the H₂ molecule from each of two complexes and subsequent dimerization through one linear oriented N₂ molecule yielding [Ru(H)₂(PXP)–N=N–Ru(H)₂(PXP)] to be thermodynamically possible. Experimentally we observed the formation of clusters of an as-yet unknown structure upon pressurizing a toluene solution of [Ru(H)₂(H₂)(PXP)] with N₂.
- [20] a) It is interesting that for complex *trans*-**Ru-4a** (that is, a pyridine ligand system), which is not capable of performing the initial hydrogen transfer to the coordinated N₂ molecule, a catalytic cycle that is otherwise the same can be calculated as for the γ -pyran and the furan system with slightly to significantly higher activation barriers in most cases. Moreover, there is an important difference in this cycle: the formation of the undesired metal nitride is possible, indicating that for this class of complexes and for the desired reaction nitrogen ligands seem to be disfavored when one tries to avoid nitride formation; b) Comparisons of the complex structures with ligand systems **a**, **b**, and **c** for the initial hydrogen transfer to coordinated N₂ (and also for other steps of the cycle) make it immediately clear that the distance between the ligand donor element (N, O) and the metal center must have implications for the bonding between donor and metal (see the Supporting Information for structures). Also, the furan ring of ligand system **c**, in contrast to the γ -pyran ring of ligand system **b**, shows a pronounced ability to rotate away from or toward the metal center, depending on the electronic situation at the metal. A detailed analysis of the structures and the chemical bonding in these complexes will be reported elsewhere.
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